

Constitutions of Rubrenolide and Rubrynlolide: an Alkene-Alkyne Pair from *Nectandra rubra*

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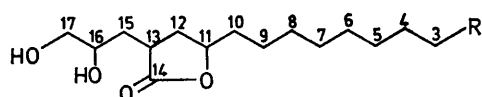
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Summary Constitutions are proposed for rubrenolide (Ia) and rubrynlolide (Ib), two structurally novel natural products isolated from *Nectandra rubra* (Lauraceae).

THE trunk wood of the Amazonian tree, *Nectandra rubra* (Mez) C. K. Allen¹ of the Lauraceae family has yielded rubrenolide and rubrynlolide which represent a novel natural product type.

Rubrenolide (Ia), C₁₇H₃₀O₄, m.p. 100°, [α]_D + 21°, ν_{CO} 1745 cm⁻¹, was identified as a terminal alkene which, on the basis of i.r. and n.m.r. evidence, is associated with a -CH₂-CH=CH₂ group. Rubrynlolide (Ib), C₁₇H₂₈O₄, m.p. 88°, [α]_D + 21°, ν_{CO} 1745 cm⁻¹, was similarly shown to be the corresponding terminal alkyne containing a -CH₂-C≡CH group. On catalytic hydrogenation rubrenolide and rubrynlolide each gave dihydrorubrenolide (Ic), C₁₇H₃₂O₄, m.p. 106–107°, [α]_D + 22°. The two hydroxy-groups present in rubrenolide were shown, using trichloroacetyl isocyanate,² to be primary and secondary, and their vicinal relationship was established by periodate cleavage of dihydrorubrenolide which gave formaldehyde and an aldehyde, C₁₆H₂₈O₃, m.p. 50–53° (ν_{CO} 1760 and 1720 cm⁻¹). Intramolecular hydrogen bonding made a direct deduction of the lactone ring size in rubrenolide uncertain (ν_{CO} 1745 cm⁻¹), but the acetonide, C₂₀H₃₆O₄, m.p. 47–48°, prepared from dihydrorubrenolide, acetone, and anhydrous copper sulphate was clearly a γ-lactone (ν_{CO} 1770 cm⁻¹). The C₁₇-skeleton of rubrenolide was determined by transforming dihydrorubrenolide to 4-methylhexadecane (III) by the sequence: (i) reduction (LiAlH₄), (ii) formation of a tetramesylate, and (iii) hydrogenolysis (LiAlH₄). Vigorous oxidation of rubrenolide with chromium trioxide in boiling acetic acid gave sebacic acid, HO₂C[CH₂]₈CO₂H. Thus, two possible constitutions (IIa) or (IIb) could be considered for rubrenolide, and a decision between these was possible on the following evidence. Mild oxidation of rubrenolide (CrO₃-HOAc; room temp.) yielded a vinylic (ν_{C=C} 1645, 990, and 910 cm⁻¹) lactonic acid (ν_{CO} 1760 and 1700 cm⁻¹), C₁₆H₂₆O₄, m.p. 95–98° (IIc), whose n.m.r. spectrum showed an ABX system (H_A τ 7.05, H_B τ 7.49, H_X m, τ ca. 7.0; J_{AB} 17.5, J_{AX} 3.5, and J_{BX} 8.5 Hz) which had to be associated with the oxidatively generated carboxy-group (HO₂C-CH_AH_B-CH_X<). A corresponding ABX system was also identifiable in the n.m.r. spectrum of the lactonic acid methyl ester (IIc).

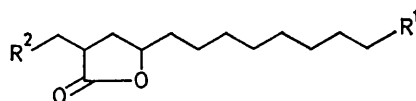
These results established the constitutions of rubrenolide (Ia) and rubrynlolide (Ib) and these proposals were fully confirmed by examination of the high resolution mass spectra and n.m.r. spectra of the two compounds and their derivatives. The assignment (Table) of seven 1H multiplets observed in the 220 MHz n.m.r. spectrum (CD₃OD-D₂O) of dihydrorubrenolide (Ic) was made possible by spin-decoupling experiments.



(Ia) Rubrenolide R = CH₂=CH

(Ib) Rubrynlolide R = CH≡C

(Ic) Dihydrorubrenolide R = CH₃-CH₂

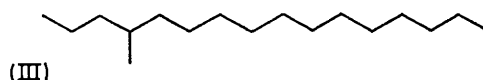


(IIa) R¹ = CH₂=CH ; R² = HO-CH₂-CH(OH)

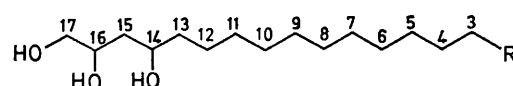
(IIb) R¹ = HO-CH₂-CH(OH) ; R² = CH₂=CH

(IIc) R¹ = CH₂=CH ; R² = CO₂H

(IId) R¹ = CH₂=CH ; R² = CO₂Me

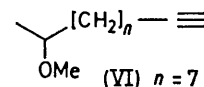
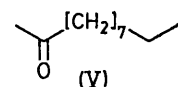


(III)



(IVa) Avocadene R = CH₂=CH

(IVb) Avocadyne R = HC≡C



(VII) n = 9, probable structure

N.m.r. spectrum (220 MHz) of dihydrorubrenolide (Ic)

τ	Multiplicity	Assignment
8.02	Octet, J 14.5, 10, 5 Hz	H _A at C-15 ^a
7.45	Octet, J 12.5, 8.5, 5.5 Hz	H _A at C-12 ^{a,b}
7.06	Complex	H at C-13
6.51	Double d, J 11, 6.5 Hz	H _B at C-17
6.38	Double d, J 11, 3.5 Hz	H _A at C-17
6.28	Complex	H at C-16
5.54	Complex	H at C-11

^a H_B at C-15 and H_B at C-12 are obscured by the [CH₂]₈ signals which extend down to τ 8.2. ^b Double irradiation at τ 5.54 produced only one clearly observable change in the spectrum: the signal at τ 7.45 changed to a double d (J 12.5, 8.5 Hz).

The branched chained C₁₇-compounds, rubrenolide (Ia) and rubrynlolide (Ib) from *Nectandra rubra* (Lauraceae) show an interesting biogenetic relation to the family of structurally related linear C₁₇-compounds [*e.g.* (IVa) and (IVb)] recently isolated from the seed of the avocado pear, *Persea americana* (Lauraceae).^{3,4} A proposed phytochemical relation^{4,5} between these two classes of C₁₇-compounds isolated from Lauraceae may be extended to include some

C₁₁- and C₁₃-compounds, (V)—(VII) from *Litsea odifera*.⁶ The natural products (Ia), (IIa), (IVa), (IVb) and related compounds⁴, (V), (VI), and (VII) may all be derived as biosynthetic variants upon the acetate-malonate route to saturated fatty acids;⁷ the biogenetic correlation between the natural product types (I) and (IV) is indicated by numbers.

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¹ C. K. Allen, *Mem. N.Y. Bot. Garden*, 1964, **10**(5), 120.

² I. R. Trehan, C. Monder, and A. K. Bose, *Tetrahedron Letters*, 1968, 67.

³ Y. Kashman, I. Néeman, and A. Lifshitz, *Israel J. Chem.*, 1969, **7**, 173; *Tetrahedron*, 1969, **25**, 4617; 1970, **26**, 1943; *Applied Microbiology*, 1970, **19**, 470; S. Bittner, S. Gazit, and A. Blumenfeld, *Phytochemistry*, 1971, **10**, 1417.

⁴ H. Magalhães Alves, D. T. Coxon, C. P. Falshaw, W. O. Godtfredsen, and W. D. Ollis, Lecture at the Symposium on Natural Products of Tropical America, Academia Brasileira de Ciências, Rio de Janeiro, August 1969; *Anais Acad. brasil. Ciênc.*, 1970, **42** (supplement), 45.

⁵ O. R. Gottlieb, in 'Chemosystematics of the Lauraceae,' Lecture at the 161st National Meeting of the American Chemical Society, Los Angeles, April 2nd, 1971, *Phytochemistry*, 1972, **11**, 1537.

⁶ W. S. Matthews, G. B. Pickering, and A. T. Umoh, *Chem. and Ind.*, 1963, 122.

⁷ F. Lynen, *Angew. Chem.*, 1965, **77**, 929; U. Henning, *Angew. Chem. Internat. Edn.*, 1966, **5**, 785; A. T. James, *Chem. in Britain*, 1968, **4**, 484 and references cited therein.